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thermal reservoir at constant temperature T_h , converts an amount W of this heat into work, and expels the remainder Q_c into another thermal reservoir at a constant temperature T_c . The efficiency e of the heat engine is the fraction of Q_h that is converted into work, so $e = W/Q_h = Q_h - Q_c / Q_h = 1 - Q_c / Q_h$ (1) Since the absorbed heat Q

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Schroeder's book is great, but the current printings (yes, the main hardcover edition, ISBN 9780201380279) are suffering from sub-par printing quality. Older printings used a high quality glossy paper. The current printings use a thicker matte paper, adding bulk to the book as a whole.

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This agrees with Schroeder's Figure 1.14. To put in some numbers, we can use the data from the earlier problem. For water, the values given by Schroeder are (at 25 C =298 K): $\rho = 2:57 \cdot 10^4 \text{ kg m}^{-3}$ (20) $T = 4:52 \cdot 10^{10} \text{ Pa}$ (21) For one mole of water, the volume is $18:068 \cdot 10^{-6} \text{ m}^3$, so $C_P C_V = (298) 18:068 \cdot 10^{-6} 2:57 \cdot 10^4 2 4:52 \cdot 10^{10} (22) = 0:787 \text{ J K}^{-1}$ (23) C

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This is a textbook for the standard undergraduate-level course in thermal physics. The book explores applications to engineering, chemistry, biology, geology, atmospheric science, astrophysics, cosmology, and everyday life.

Thermal physics deals with collections of large numbers of particles - typically 10^{23} or so. Examples include the air in a balloon, the water in a lake, the electrons in a chunk of metal, and the photons given off by the sun. We can't possibly follow every detail of the motions of so many particles. So in thermal physics we assume that these motions are

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random, and we use the laws of probability to predict how the material as a whole ought to behave. Alternatively, we can measure the bulk properties of a material, and from these infer something about the particles it is made of. This book will give you a working understanding of thermal physics, assuming that you have already studied introductory physics and calculus. You will learn to apply the general laws of energy and entropy to engines, refrigerators, chemical reactions, phase transformations, and mixtures. You will also learn to use basic quantum physics and powerful statistical methods to predict in detail how temperature affects molecular speeds, vibrations of solids, electrical and magnetic behaviors, emission of light, and exotic low-temperature phenomena. The problems and worked examples explore applications not just within physics but also to engineering, chemistry, biology, geology, atmospheric science, astrophysics, cosmology, and everyday life.

This fully updated and expanded new edition continues to provide the most readable, concise, and easy-to-follow introduction to thermal physics. While maintaining the style of the original work, the book now covers statistical mechanics and incorporates worked examples systematically throughout the text. It also includes more problems and essential updates, such as discussions on superconductivity, magnetism, Bose-Einstein condensation, and climate change. Anyone needing to acquire an intuitive understanding of thermodynamics from first principles will find this third edition indispensable. Andrew Rex is professor of physics at the University of Puget Sound in Tacoma, Washington. He is author of several textbooks and the

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popular science book, Commonly Asked Questions in Physics.

CONGRATULATIONS TO HERBERT KROEMER, 2000 NOBEL LAUREATE FOR PHYSICS
For upper-division courses in thermodynamics or statistical mechanics, Kittel and Kroemer offers a modern approach to thermal physics that is based on the idea that all physical systems can be described in terms of their discrete quantum states, rather than drawing on 19th-century classical mechanics concepts.

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themselves when faced with unfamiliar situations.

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